Exp. Charles' Law

1. What is the gas used in today's experiment?

2. If the temperature of a given volume of gas is doubled while the pressure remains constant, how will the volume change?

3. At what temperatures will the volume of the gas in this experiment be measured?

4. Write the mathematical equation for Charles' Law.

5. Which of the following represents Charles' Law?

6. List 2 reasons for why the flask needs to be raised in the ice bath until the water level inside the flask is at the same height as the water level in the bath.

7. Why should the pinch clamp be open during the heating of the flask?

8. If you would like to design an experiment to verify Charles' Law, what are the two temperatures that seem practical to use? Why?

9. Give an everyday example of Charles' law.

10. Why should you dry the flask thoroughly before doing the Charles' law experiment?